

Name \_\_\_\_\_ Period \_\_\_\_\_ Date \_\_\_\_\_

### Fill In Table Using the Periodic Table (#1)

The **atomic mass** is the average mass of an element (given as a decimal on the periodic table.) Atomic mass = protons + neutrons (The mass of an atom comes from the nucleus)

The **atomic number** (whole number in block of Periodic Table) = **# of protons** (p+)

Consider elements to be neutral in charge - the **number of electrons** will be the **same as the number of protons**.

The **number of neutrons** = **Atomic mass** - **number of protons**

An atom of an element is considered an **isotope** when there are **more neutrons than protons**. (You must understand that even if you calculate the number of neutrons to be greater than the number of protons for a given element using the formula above, this does not mean that all atoms of that element are isotopes.)

*Fill in the missing information in the table below using your periodic table.*

Name of Element	Sym bol	Atomic mass	Atomic Number (# of p+)	# of e-	# of Neutrons	Isotope?
	K					
	Ca					
	Sn					
	Pb					
	Al					
	Kr					
	B					
	N					
	U					
	Pu					
	Xe					
	Ce					
	Fr					
	Ra					
	Ti					